CAREER ENDEAVOUR
Best Institute for IIT-JAM, NET & GATE

NEET 2024

CODE: T4

SECTION-A (Q.1 TO Q.35)

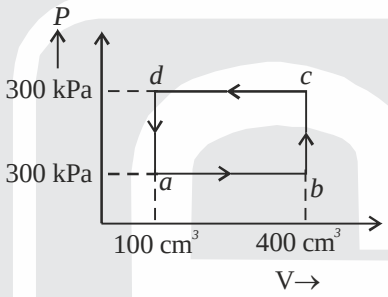
1. The moment of inertia of a thin rod about an axis 6 passing through its mid point and perpendicular to the rod is 2400 g cm². The length of the 400 g rod is nearly

(a) 20.7 cm (b) 72.0 cm (c) 8.5 cm (d) 17.5 cm

2. A bob is whirled in a horizontal plane by means of a string with an initial speed of ω rpm. The tension in the string is T. If speed becomes 2ω while keeping the same radius, the tension in the string becomes

(a) $\frac{T}{4}$ (b) $\sqrt{2}T$ (c) T (d) $4T$

3. A thermodynamic system is taken through the cycle abcda. The work done by the gas along the path be is



(a) –90 J (b) –60 J (c) Zero (d) 30 J

4. $^{290}_{82}X \xrightarrow{a} Y \xrightarrow{e^+} Z \xrightarrow{\beta^-} P \xrightarrow{e^-} Q$

In the nuclear emission stated above, the mass number and atomic number of the product Q respectively, are

(a) 288, 82 (b) 286, 81 (c) 280, 81 (d) 286, 80

5. An unpolarised light beam strikes a glass surface at Brewster's angle. Then

(a) both the reflected and refracted light will be completely polarised.

(b) the reflected light will be completely polarised but the refracted light will be partially polarised.

(c) the reflected light will be partially polarised.

(d) the refracted light will be completely polarised.

6. If c is the velocity of light in free space, the correct statements about photon among the following are:

A. The energy of a photon is $E = h\nu$.

B. The velocity of a photon is c.

C. The momentum of a photon $p = h\nu/c$

D. In a photon-electron collision, both total energy and total momentum are conserved.

E. Photon possesses positive charge.

Choose the correct answer from the options given below:

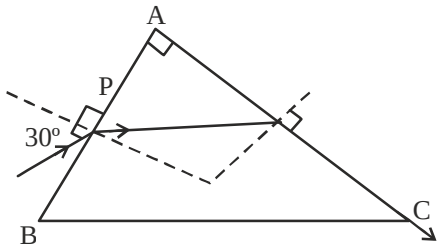
(a) A, C and D only (b) A, B, D and E only

(c) A and B only (d) A, B, C and D only

7. Two bodies A and B of same mass undergo completely inelastic one dimensional collision. The body A moves with velocity v_1 while body B is at rest before collision. The velocity of the system after collision is v_2. The ratio $v_1 : v_2$ is:

(a) 4:1 (b) 1:4 (c) 1:2 (d) 2:14

8. A light ray enters through a right angled prism at point P with the angle of incidence 30° as shown in figure. It travels through the prism parallel to its base BC and emerges along the face AC. The refractive index of the prism is:



(a) $\frac{\sqrt{3}}{4}$ (b) $\frac{\sqrt{3}}{2}$ (c) $\frac{\sqrt{5}}{4}$ (d) $\frac{\sqrt{5}}{2}$

9. If $x = 5\sin\left(\pi t + \frac{\pi}{3}\right)$ represents the motion of a particle executing simple harmonic motion, the amplitude and time period of motion, respectively, are:

(a) 5 cm, 1 s (b) 5m, 1s (c) 5 cm, 2s (d) 5m, 2s

10. At any instant of time *t*, the displacement of any particle is given by 2*t*-1 (Sl unit) under the influence of force of 5N. The value of instantaneous power is (in SI unit):

(a) 7 (b) 6 (c) 10 (d) 5

11. A tightly wound 100 turns coil of radius 10 cm carries a current of 7 A. The magnitude of the magnetic field at the centre of the coil is (Take 15 permeability of free space as $4\pi \times 10^{-7}$ SI units):

(a) 4.4 mT (b) 44 T (c) 44 mT (d) 4.4 T

12. A particle moving with uniform speed in a circular path maintains:

(a) constant velocity and varying acceleration

(b) varying velocity and varying acceleration

(c) constant velocity

(d) constant acceleration

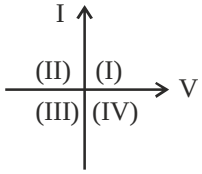
13. A logic circuit provides the output Y as per the follow ing truth table

A	B	Y
0	0	1
0	1	0
1	0	1
1	1	0

The expression for the output Y is

(a) $\bar{B}$ (b) B (c) $AB + \bar{A}$ (d) $A \cdot \bar{B} + \bar{A}$

14. Consider the following statements A and B and identify the correct answer:



(A)For a solar-colt the I-V characteristics lies in the IV quadrant of the given graph.

(B) In a reverse blased *pn* junction diode, the current measured in (μA), is due to majority charge carriers.

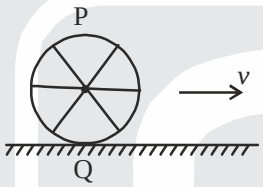
(a) Both A and B are correct. (b) Both A and Bare incorrect.

(c) A is correct bud B is incorrect. (d) A is incorrect but B is correct.

15. In an ideal transformer, the turns ratio is $\frac{N_p}{N_s} = \frac{1}{2}$. The ratio $V_s : V_p$ is equal to (the symbols carry their usual meaning)

(a) 1: 1 (b) 1 : 4 (c) 1 : 2 (d) 2 : 1

16. A wheel of a bullgok cart is rolling on a level road as shown in the figure below. If its linear speed is v in the direction shown, which one of the following optidis is correct (P and Q are any highest and lowest points on the wheel, respectively)?



(a) Both the points P and Q move with equal speed.

(b) Point P has zero speed.

(c) Point P moves slower than point Q

(d) Point P moves faster than point (Q)

17. If the monochromatic source in Young's double slit experiment is replaced by white light, then

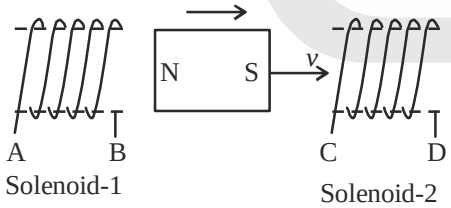
(a) there will be a central bright white fringe surrounded by a few coloured fringes.

(b) all bright fringes will be of equal width.

(c) interference pattern will disappear.

(d) there will be a central dark fringe surrounded by a few coloured fringes.

18.



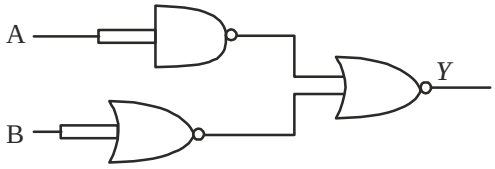
In the above diagram, a strong bar magnet is moving towards solenoid-2 from solenoid-1. The direction of induced current in solenoid-1 and that in solenoid-2, respectively, are through the directions:

(a) AB and CD (b) BA and DC (c) AB and DC (d) B4 and CD

19. In a vernier calipers. (N+1) divisions of vernier seale coincide with N divisions of main scale. If I MSD represents 0.1 mm, the vernier constant (in cm) is:

(a) 100 N (b) 10 (N + 1) (c) 1/10 N (d) 1/100(N+1)

20. The output (Y) of the given logic gate is similar to the output of an/a



(a) OR (b) AND gate (c) NAND gate (d) NOR gate

21. Given below are two statements: one is labelled as **Assertion A** and the other is labelled as **Reason R**.

Assertion A: The potential (I) at any axial point at 2 m distance (r) from the centre of the dipole of dipole moment vector $\vec{P}$ of magnitude, $4\times10^{-6}C\,m$, is $\pm9\times10^{3}V$ (Take, $\frac{1}{4\pi\in_0}=9\times10^{9}$ SI units)

Reason: $v=\pm\frac{2p}{4\pi\in_0 r^2}$, where r is the distance of any axial point, situated at 2m from the centre of the dipole.

In the light of the above statements, choose the correct answer from the options given below:

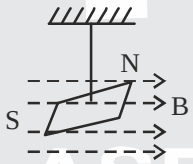
(a) A is true but R is false.

(b) A is false but R is true.

(c) Both A and R are true and R is the correct explanation of A.

(d) Both A and R are true and R is NOT the correct explanation of A

22. In a uniform magnetic field of 0.049 T, a magnetic needle performs 20 complete oscillations in 5 seconds as shown. The moment of inertia of the needle is 9.8 10^6 kg m². If the magnitude of magnetic moment of the needle is $x\times10^{-5}$ Am²; then the value of 'x' is:



(a) $50\,\pi^2$ (b) $1280\,\pi^2$ (c) $5\,\pi^2$ (d) $128\,\pi^2$

23. Match **List-I** with **List-II**.

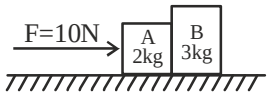
List-I (Material)	**List-II (Susceptibility (χ))**
(A) Diamagnetic	I. $\chi=0$
(B) Ferromagnetic	II. $0>\chi\geq-1$
(C) Paramagnetic	III. $\chi>>1$
(D) Non-magnetic	IV. $0<\chi<\varepsilon$ (a small positive number)

Choose the correct answer from the options given below

(a) A-III, B-II, C-I, D-IV (b) A-IV, B-III, C-II, D-I

(c) A-II, B-III, C-IV, D-I (d) A-II, B-I, C-III, D-IV

24. A horizontal force 10 N is applied to a block 1 as shown in figure. The mass of blocks A and B are 2 kg and 3 kg, respectively. The blocks slide over a frictionless surface. The force exerted by block A on block B is:



(a) 6N (b) 10 N (c) Zero (d) 4 N

25. Given below are two statements:

Statement I: Atoms are electrically neutral as they contain equal number of positive and negative charges.
Statement II: Atoms of each element are stable

and emit their characteristic spectrum. In the light of the above statements, choose the most appropriate answer from the options given below:

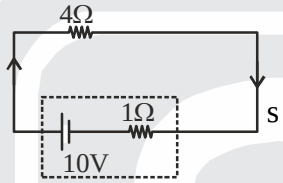
(a) Statement I is correct but Statement II is incorrect.

(b) Statement I is incorrect but Statement II is correct.

(c) Both Statement I and Statement II are correct.

(d) Both Statement I and Statement Il are incorrect.

26. The terminal voltage of the battery, whose emf is 10 and internal resistance 1Ω, when connected through an external resistance of 4Ω as shown in the figure is:



(a) 8V (b) 10V (c) 4V (d) 6 V

27. A wire of length '*l*' and resistance 100 Ω is divided into 10 equal parts. The first 5 parts are connected in series while the next 5 parts are connected in parallel. The two combinations are again connected in series. The resistance of this final combination is

(a) 55 Ω (b) 60 Ω (c) 26 Ω (d) 52 Ω

28. The maximum elongation of a steel wire of 1 m length if the elastic limit of steel and its Young's modulus, respectively, are 8×10^{8} N m^{-2}, and 2×10^{11} N m^{-2} is:

(a) 40 mm (b) 8 mm (c) 4 mm (d) 0.4 mm

29. A thin flat circular disc of radius 4.5 cm is placed gently over the surface of water. If surface tension of water is 0.07 Nm, then the excess force required to take it away from the surface is:

(a) 1.98 mN (b) 99 N (c) 19.8 mN (d) 198 N

30. Match List I with List II.

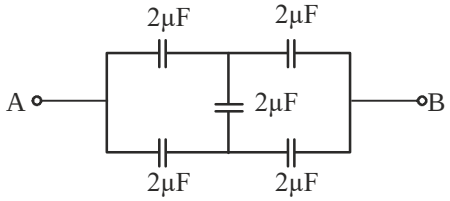
List-I (Spectral Lines of Hydrogen for transitions from)	**List-II (Wavelengths (nm))**
(A) $n_2 = 3$ to $n_1 = 2$	I. 410.2
(B) $n_2 = 4$ to $n_1 = 2$	II. 434.1
(C) $n_2 = 5$ to $n_1 = 2$	III. 656.3
(D) $n_2 = 6$ to $n_1 = 2$	IV. 486.1

Choose the correct answer from the options given below

(a) A-IV, B-III, C-I, D-II (b) A-I, B-II, C-III, D-IV

(c) A-II, B-I, C-IV, D-III (d) A-III, B-IV, C-II, D-I

31. In the following circuit, the equivalent capacitance between terminal 4 and terminal B is

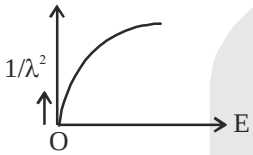


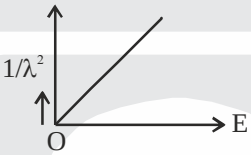
(a) $0.5\ \mu F$ (b) $4\ \mu F$ (c) $2\ \mu F$ (d) $1\ \mu F$

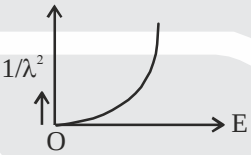
32. The mass of a planet is $\frac{1}{10}^{th}$ that of the earth and its diameter is half that of the earth. The acceleration due to gravity on that planet is:

(a) 4.9 ms^{-2} (b) 3.92 ms^{-2} (c) 19.6 ms^{-2} (d) 9.8 ms^{-2}

33. The graph which shows the variation of $\left(\frac{1}{\lambda^2}\right)$ and its kinetic energy E is (where λ is de Broglie wavelength of a free particle):

(a) 

(b) 

(c) 

(d) 

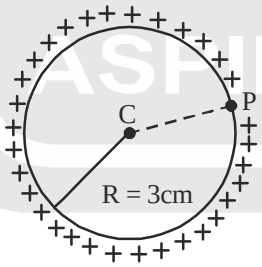
34. The quantities which have the same dimensions as those of solid angle, are:

(a) strain and arc (b) angular speed and stress

(c) strain and angle (d) stress and angle

35. A thin spherical shell is charged by some source. The potential difference between the two points C and P (in) shown in the figure is:

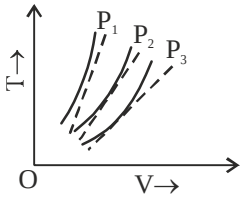
(Take: $\frac{1}{4\pi \in_0} = 9\times10^9$ (SI units)



(a) 0.5×10^5 (b) Zero (c) 3×10^5 (d) 1×10^5

SECTION-B (Q.36 TO Q.50)

36. The following graph represents the T-V curves of an ideal gas (where T is the temperature and the volume) at three pressures P_1, P_2 and P_3 compared with those of Charles's law represented as dotted lines



Then the correct relation is:

(a) $P_2 > P_1 > P_3$ (b) $P_1 > P_2 > P_3$ (c) $P_3 > P_2 > P_1$ (d) $P_1 > P_3 > P_2$

37. The property which is not of an electromagnetic wave travelling in free space is that:

(a) they travel With a speed equal to $\frac{1}{\sqrt{\mu_0 \in_0}}$

(b) they originate from charges moving with uniform speed.

(c) they are transverse in nature.

(d) the energy density in electric field is equal to energy density in magnetic field.

38. A small telescope has an objective of focal length 140 cm and an eye piece of focal length 5.0 cm. The magnifying power of telescope for viewing a distant object is:

(a) 17 (b) 32 (c) 34 (d) 28

39. A parallel plate capacitor is charged by connecting it to a battery through a resistor. If 1 is the current in the circuit, then in the gap between the plates:

(a) displacement current of magnitude equal to I flows in a direction opposite to that of I.

(b) displacement current of magnitude greater than I flows but can be in any direction.

(c) there is no-current.

(d) displacement current of magnitude equal to I flows in the same direction as I

40. A metallic bar of Young's modulus, 0.5x1011 N m-2 and coefficient of linear thermal expansion 10^{-5} °C^{-1}, length 1 m and area of cross-section 10-3 m² is heated from 0°C to 100°C without expansion or bending. The compressive force developed in it is:

(a) 100 $\times 10^3$ N (b) 2×10^3 N (c) 5×10^3 N (d) 50×10^3 N

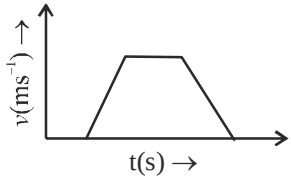
41. Two heaters A and B have power rating of 1 kW and 2 kW, respectively. Those two are first connected in series and then in parallel to a fixed power source. The ratio of power outputs for these two cases is:

(a) 1:2 (b) 2:3 (c) 1:1 (d) 2:9

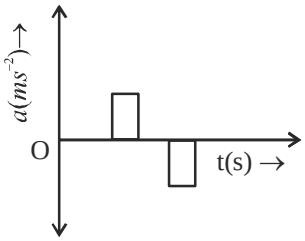
42. An iron bar of length L has magnetic moment M. It is bent at the middle of its length such that the two arms make an angle 60° with each other. The magnetic moment of this new magnet is:

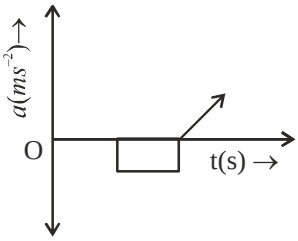
(a) $2M$ (b) $M/\sqrt{3}$ (c) M (d) $M/2$

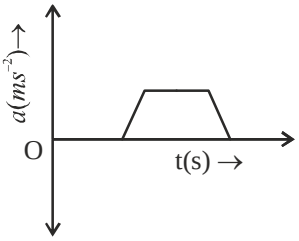
43. The velocity (v)-time (t) plot of the motion of a body is shown below:

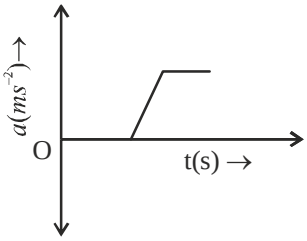


The acceleration (a)-time (t) graph that best suits this motion is:

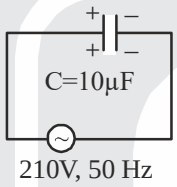
(a) 

(b) 

(c) 

(d) 

44. A 10 μF capacitor is connected to a 210 V, 50 Hz source as shown in figure. The peak current in the circuit is nearly $(\pi = 3.14)$:

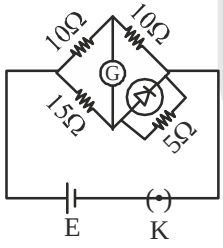


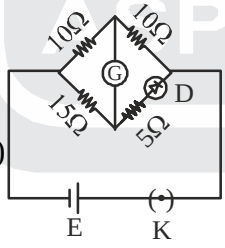
(a) 1.20 A (b) 0.35 A (c) 0.58 A (d) 0.93 A

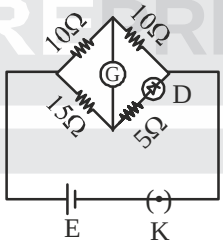
45. A force defined by $F = \alpha t^2 + \beta t$ acts on a particle at a given time t. The factor which is dimensionless, if α and β are constants, is:

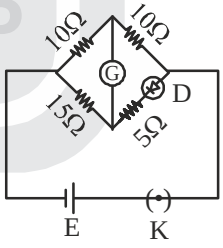
(a) $\alpha\beta t$ (b) $\alpha\beta / t$ (c) $\beta t / \alpha$ (d) $\alpha t / \beta$

46. Choose the correct circuit which can achieve the bridge balance

(a) 

(b) 

(c) 

(d) 

47. If the mass of the bob in a simple pendulum is increased to thrice its original mass and its length is made half its original length, then the new time period of oscillation is $x/2$ times its original time period. Then the value of x is:

(a) $2\sqrt{3}$ (b) 4 (c) $\sqrt{3}$ (d) $\sqrt{2}$

48. If the plates of a parallel plate capacitor connected to a battery are moved close to each other, then
(A) the charge stored in it. Increases.
(B) the energy stored in it. decreases.
(C) its capacitance increases.
(D) the ratio of charge to its potential remains the same.
(E) the product of charge and voltage increases.
Choose the most appropriate answer from the options given below:
(a) B, D and E only (b) A, B and C only
(c) A. B and E only (d) A, C and E only

49. The minimum energy required to launch a satellite of mass m from the surface of earth of mass M and radius R in a circular orbit at an altitude of 2R from the surface of the earth is

(a) $\frac{GmM}{2R}$ (b) $\frac{GmM}{3R}$ (c) $\frac{5GmM}{6R}$ (d) $\frac{2GmM}{3R}$

50. A sheet is placed on a horizontal surface in front of a strong magnetic pole. A force is needed to :
A. hold the sheet there if it is magnetic.
B. hold the sheet there if it is non-magnetic.
C. move the sheet away from the pole with uniform velocity if it is conducting.
D. move the sheet away from the pole with uniform velocity if it is both, non-conducting and non-polar.
Choose the correct statement(s) from the options given below:
(a) A, C and D only (b) C only (c) B and D only (d) A and C only

SECTION-B (Q.51 TO Q.85)

51. Match List I with List II.

List I (Process)	List II (Conditions)
1. No heat exchange	A. Isothermal process
II. carried out at constant temperature	B. Isochoric process
III. Carried out at constant volume	C. Isobaric process
IV. Carried out at constant pressure	D. Adiabatic process

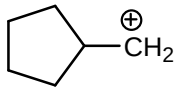
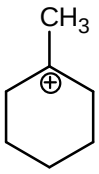
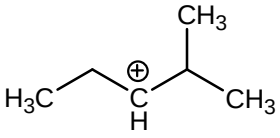
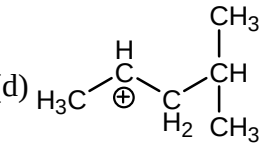
Choose the correct answer from the options given below:
(a) A-I, B-II, C-III, D-IV (b) A-II, B-III, C-IV, D
(c) A-IV, B-III, C-II, DI (d) A-IV, B-II, C-III, D-I

52. Match **List -I** with **List-II**

List-I (Complex)	**List-II (Type of isomerism)**
(A) $\left[Co(NH_3)_5(NO_2)\right]Cl_2$	(I) Solvate isomerism
(B) $\left[Co(NH_3)_5(SO_4)\right]Br$	(II) Linkage isomerism
(C) $\left[Co(NH_3)_6\right]\left[Cr(CN)_6\right]$	(III) Ionization isomerism
(D) $\left[Co(H_2O)_6\right]Cl_3$	(IV) Coordination isomerism

Choose the correct answer from the options given below
(a) A-I, B-IV, C-III, D-II (b) A-II, B-IV, C-III, D-I
(c) A-II, B-III, C-IV, D-I (d) A-I, B-III, C-IV, D-II

53. The most stable carbocation among the follwoing is

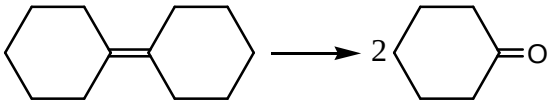
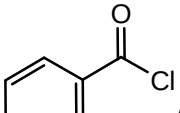
(a)  (b)  (c)  (d) 

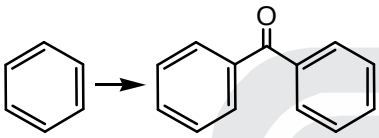
54. On heating, some solid substances change from solid to vapour state without passing through liquid state. The technique used for the purification of such solid substances based on the above principle is known as

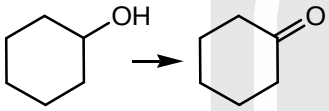
(a) Distillation (b) Chromatography (c) Crystallization (d) Sublimation

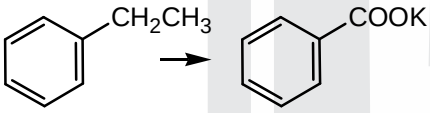
55. Match the **List-I** with **List-II**

List-I (Reaction) **List-II (reagents/Condition)**

(A)  (I)  /Anhyd. $AlCl_3$

(B)  (II) CrO_3

(C)  (III) $KMnO_4/KOH, \Delta$

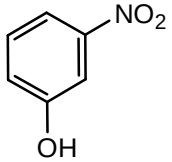
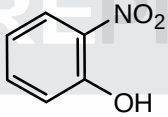
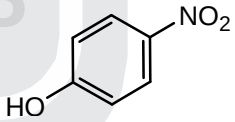
(D)  (IV) (i) O_3; (ii) Zn-H_2O

Choose the correct answer from the options given below

(a) A-IV, B-I, C-II, D-III (b) A-I, B-IV, C-II, D-III

(c) A-IV, B-I, C-III, D-II (d) A-III, B-I, C-II, D-IV

56. Intramolecular hydrogen bonding is present in

(a)  (b) HF (c)  (d) 

57. The highest nümber of helium atoms is in

(a) 4 g of helium (b) 2.271098 L of helium at STP

(c) 4 mol of helium (d) 4 u of helium

58. For the reaction $2A \rightleftharpoons B + C$, $K_c = 4\times10^{-3}$. At a given time, the composition of reaction mixture is:

$[A] = [B] = [C] = 2\times10^{-3}M$

Then, which of the following is correct?

(a) Reaction has a tendency to go in backward direction.

(b) Reaction has gone to completion in forward direction

(c) Reaction is at equilibrium.

(d) Reaction has a tendency to go in forward direction

59. The E^o value for the Mn^{3+}/Mn^{2+} couple is more positive than that of Cr^{3+}/Cr^{2} or Fe^{3+}/Fe^{2+} due to change of
(a) d^4 to d^5 configuration (b) d^3 to d^5 configuration
(c) d^5 to d^4 configuration (d) d^5 to d^2 configuration

60. Fehling's solution 'A' is
(a) alkaline solution of sodium potassium tartrate (Rochelle's salt)
(b) aqueous sodium citrate
(c) aqueous copper sulphate
(d) alkaline copper sulphate

61. Match List I with List II.

List I (Compound)	**List II** (Shape/geometry)
(A) NH_3	(I) Trigonal Pyramidal
(B) BrF_5	(II) Square Planar
(C) XeF_4	(III) Octahedral
(D) SF_6	(IV) Square Pyramidal

Choose the correct answer from the options given below:
(a) A-III, B-IV, C-1, D-II (b) A-II, B-III, C-IV, D-I
(c) A-I, B-IV, C-II, D-III (d) A-II, B-IV, C-III, D-I

62. Given below are two statements:

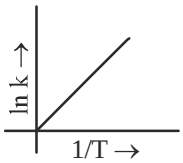
Statement I: Both $\left[Co(NH_3)_6\right]^{3+}$ and $\left[CoF_6\right]^{3-}$ complexes are octahedral but differ in their magnetic behaviour.

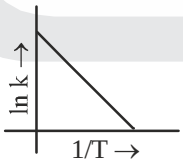
Statement II: $\left[Co(NH_3)_6\right]^{3+}$ is diamagnetic whereas $\left[CoF_6\right]^{3-}$ is paramagnetic.

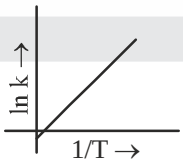
In the light of the above statements, choose the ***correct*** answer from the options given below:
(a) Statement I is true but Statement II is false.
(b) Statement I is false but Statement II is true.
(c) Both Statement I and Statement II are true.
(d) Both Statement I and Statement II are false.

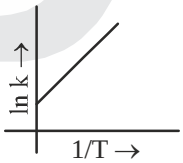
63. Among Group 16 elements, which one does **NOT** show-2 oxidation state?
(a) Te (b) Po (c) O (d) Se

64. Which plot of ln k vs 1/T is consistent with Arrhenius equation?

(a) 

(b) 

(c) 

(d) 

65. Arrange the following elements in increasing order of electronegativity:
N, O, F, C, Si
Choose the correct answer from the options giver below:
(a) O<F<N<C<Si (b) F<O<N<C<Si
(c) Si<C<N<O<F (d) Si<C<O<N<F

66. Given below are two statements:
Statement I: The boiling point of three isomeric pentanes follows the order
n-pentane > isopentane > neopentane

Statement II: When branching increases, the molecule attains a shape of sphere. This results in smaller surface area for contact, due to which the intermolecular forces between the spherical molecules are weak, thereby lowering the boiling point.

In the light of the above statements, choose the *most appropriate answer* from the options given below:

(a) Statement I is correct but Statement II is incorrect.

(b) Statement I is incorrect but Statement II is correct.

(c) Both Statement I and Statement Il are correct.

(d) Both Statement I and Statement II are incorrect.

67. Which reaction is NOT a redox reaction?

(a) $H_2 + Cl_2 \rightarrow 2HCl$ (b) $BaCl_2 + Na_2SO_4 \rightarrow BaSO_4 + 2NaCl$

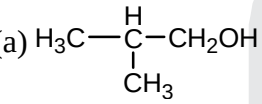
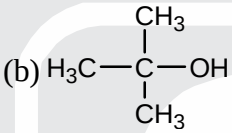
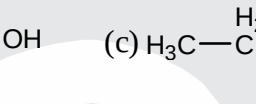
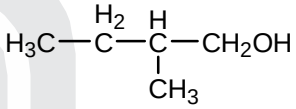
(c) $Zn + CuSO_4 \rightarrow ZnSO_4 + Cu$ (d) $2KClO_3 + I_2 \rightarrow 2KIO_3 + Cl_2$

68. Arrange the following-elements in increasing order of first ionization enthalpy: Li, Be, B, C, N. Choose the correct answer from the options given below:

(a) Li < Be < C < B < N (b) Li < Be < N < B<C

(c) Li<Be < B < C < N (d) Li < B < Be < C < N

69. Which one of the following alcohols reacts instantaneously with Lucas reagent?

(a)  (b)  (c)  (d) 

70. Match List I with List II.

List I (Molecule)	**List II (Number and types of bond/s between two carbon atoms)**
A. ethane	I. one σ-bond and two π-bonds
B. ethene	II. two π-bonds
C. carbon molecule, C_2	III. one σ-bond
D. ethyne	IV. one σ-bond and one π-bond

Choose the correct answer from the options given below:

(a) A-III, B-IV, C-II, D-I (b) A-III, B-IV, C-I, D-II

(c) A-1, B-IV, C-II, D-III (d) A-IV, B-III, C-II, D-I

71. Given below are two statements:

Statement I: The boiling point of hydrides of Group 16 clements follow the order

$H_2O > H_2Te > H_2Se > H_2S$

Statement II: On the basis of molecular mass, H_2O is expected to have lower boiling point than the other members of the group but due to the presence of extensive H-bonding in H_2O, it has higher boiling poiit.

In the light of the above statements, choose the correct answer from the options given below:

(a) Statement I is true but Statement II is false.

(b) Statement I false but Statement II is true.

(c) Both Statement I and Statement II are true.

(d) Both Statentent I and Statement II are false.

72. Given below are two statements:

Statement I: Aniline does not undergo Friedel- Crafts alkylation reaction.

Statement II: Aniline cannot be prepared through Gabriel synthesis.

In the light of the above statements, choose the correct answer from the options given below:

(a) Statement I is true but Statement II is false

(b) Statement I is false but Statement II is true.

(c) Both Statement I and Statement II are true.

(d) Both Statement I and Statement II are false.

73. Match **List I** with **List II**.

List I (Conversion)	**List II (Number of Faraday required)**
(A) 1 mol of H_2O to O_2	I. 3F
(B) 1 mol of MnO_4^{-1} to Mn^{2+}	II. 2F
(C) 1.5 mol of Ca from molten $CaCl_2$	III. 1F
(D) 1 mol of FeO to Fe_2O_3	IV. 5F

Choose the correct answer from the options given below:

(a) A-II, B-III, GI, D-IV (b) A-III, B-IV, GII, D-I

(c) A-II, B-IV, C-I, D-III (d) A-III, B-IV, C-1, D-II

74. In which of the following equilibria. K_p and K_c are **NOT** equal?

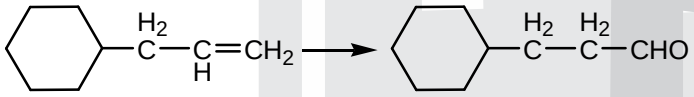
(a) $CO_{(g)} + H_2O_{(g)} \rightleftharpoons CO_{2(g)} + H_{2(g)}$ (b) $2BrCl_{(g)} \rightleftharpoons Br_{2(g)} + Cl_{2(g)}$

(c) $PCl_{5(g)} \rightleftharpoons PCl_{3(g)} + Cl_{2(g)}$ (d) $H_{2(g)} + I_{2(g)} \rightleftharpoons 2HI_{(g)}$

75. The Henry's law constant (K) values of three gases (A. B. C) in water are 145, 2×10^{-5} and 35 kbar, respectively. The solubility of these gases in water follow the order:

(a) A > C > B (b) A > B > C (c) B > A > C (d) B > C > A

76. Identify the correct reagents that would bring about the following transformation.



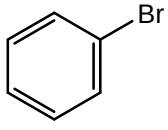
(a) (i) BH_3, (ii) H_2O_2 / OH^-, (iii) alk. $KMnO_4$, (iv) H_3O^+

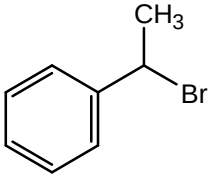
(b) (i) H_2O/H^+, (ii) PCC

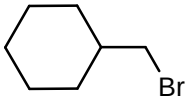
(c) (i) H_2O/H^+, (ii) CrO_3

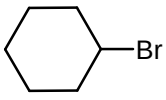
(d) (i) BH_3, (ii) H_2O_2 / OH^-, (iii) PCC

77. The compound that will undergo S_N1 reaction with the fastest rate is

(a) 

(b) 

(c) 

(d) 

78. The energy of an electron in the ground state (n=1) for He ion is-x J, then that for an electron in n=2 state for Be^{3+} ion in J is

(a) $-4x$ (b) $-\frac{4}{9}x$ (c) $-x$ (d) $-\frac{x}{9}$

79. A compound with a molecular formula of C_6H_{14} has two tertiary carbons. Its IUPAC name is:
(a) 2, 3-dimethylbutane (b) 2, 2-dimethylbutane.
(c) n-hexane (d) 2-methylpentane

80. The reagents with which glucose does not react to give the corresponding tests/products are
A. Tollen's reagent B. Schiff's reagent C. HCN
D. NH_2OH
E. $NaHSO_3$
Choose the correct options from the given below
(a) B and E (b) E and D (c) B and C (d) A and D

81. 'Spin only' magnetic moment is same for which of the following ions?
A. Ti^{3+} B. Cr^{2+} C. Mn^{2+}
D. Fe^{2+} E. Sc^{3+}
Choose the most appropriate answer from the options given below:
(a) B and C only (b) A and D only
(c) B and D only (d) A and E only

82. Match List I with List II

List-I **Quantum Number**	**List-II** **Information provided**
(A) m_ℓ	(I) shape of orbital
(B) m_s	(II) size of orbital
(C) *l*	(III) orientation of orbital
(D) *n*	(IV) orientation of spin of electron

Choose the correct answer from the options given below:
(a) A-III. B-IV, C-II, D-I (b) A-II, B-1, C-IV, D-III
(c) A-1, B-III, C-II, D-IV (d) A-III, B-IV, C-1, D-II

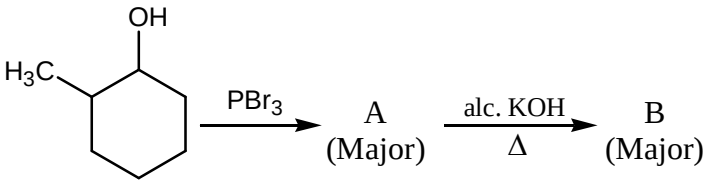
83. 1 gram of sodium hydroxide was treated with 25 mL of 0.75 M HCI solution, the mass of sodium hydroxide left unreacted is equal to
(a) Zero mg (b) 200 mg (c) 750 mg (d) 250 mg

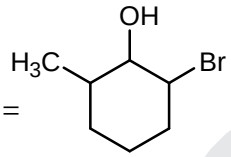
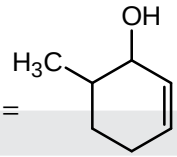
84. In which of the following processes entropy increases?
A. A liquid evaporates to vapour.
B. Temperature of a crystalline solid lowered from 130 K to 0 K.
C. $2NaHCO_{3(s)} \rightarrow Na_2CO_{3(s)} + CO_{2(s)} + H_2O_{(g)}$
D. $Cl_{2(g)} \rightarrow 2Cl_{(g)}$
Choose the correct answer from the options given below:
(a) A, C and D (b) C and D (c) A and C (d) A, B and D

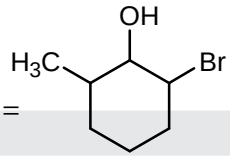
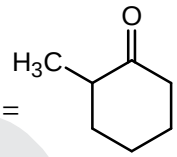
85. Activation energy of any chemical reaction can be calculated if one knows the value of
(a) orientation of reactant molecules during collision.
(b) rate constant at two different temperatures.
(c) rate constant at standard temperature.
(d) probability of collision.

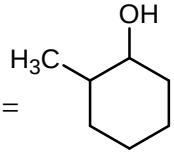
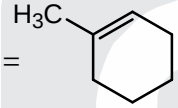
SECTION-B (Q.86 TO Q.100)

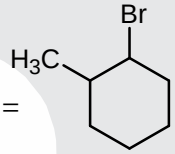
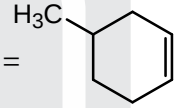
86. Major products A and B formed in the following reaction sequence, are



(a) A =  B = 

(b) A =  B = 

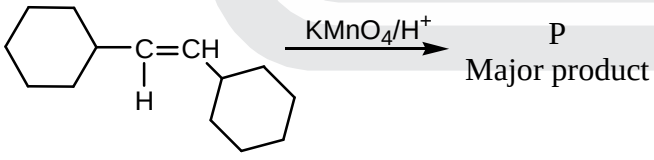
(c) A =  B = 

(d) A =  B = 

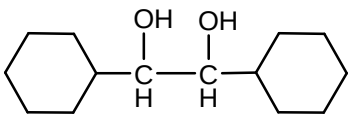
87. The work done during reversible isothermal expansion of one mole of hydrogen gas at 25°C from pressure of 20 atmosphere to 10 atmosphere is: (Given R = 2.0 cal K^{-1} mol^{-1})
(a) 413.14 calories (b) 100 calories (c) 0 calorie (d) –413.14 calories

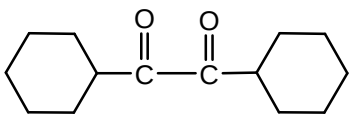
88. Consider the following reaction in a sealed vessel at equilibrium with concentrations of $N_2 = 3.0\times10^{-3}M, O_2 = 4.2\times10^{-3}$ and $NO = 2.8\times10^{-3}M$?. $2NO_{(g)} \rightleftharpoons N_{2(g)} + O_{2(g)}$. If 0.1 mol L^{-1} of $NO_{(g)}$ is taken in a closed vessel. what will be degree of dissociation (α) of $NO_{(g)}$ at equilibrium?
(a) 0.8889 (b) 0.717 (c) 0.00889 (d) 0.0889

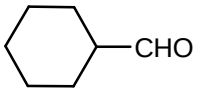
89. For the given reaction:

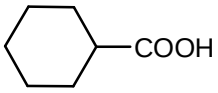


(P) is

(a) 

(b) 

(c) 

(d) 

90. The pair of lanthanoid ions which are diamagnetic is
(a) Gd^{3+} and Eu^{3+} (b) Pm^{3+} and Sm^{3+} (c) Ce^{4+} and Yb^{2+} (d) Ce^{3+} and Eu^{2+}

91. Identify the major product C formed in the following reaction sequence:

$$CH_3-CH_2-CH_2-I \xrightarrow{NaCN} (A) \xrightarrow[\text{Partial hydrolysis}]{OH^-} (B) \xrightarrow[Br_2]{NaOH} \begin{pmatrix} C \\ \text{Major} \end{pmatrix}$$

(a) butanamide (b) α-bromobutanoic acid (c) propylamine (d) butylamine

92. The products A and B obtained in the following reactions, respectively, are

$$3ROH + PCl_3 \longrightarrow 3RCl + A$$

$$ROH + PCl_5 \rightarrow RCl + HCl + B$$

(a) H_3PO_4 and $POCl_3$ (b) H_3PO_3 and $POCI_3$
(c) $POCl_3$, and H_3PO_3 (d) $POCI_3$ and H_3PO_4

93. Given below are certain cations. Using inorganic qualitative analysis, arrange them in increasing group number from 0 to VI.

A. Al^{3+} B. Cu^{2+} C. Ba^{2+}
D. Co^{2+} E. Mg^{2+}

Choose the correct answer from the options give below:

(a) E, C, D, B, A (b) E, A, B, C, D (c) B, A,D, C, E (d) B, C, A, D, E

94. A compound X contains 32% of A, 20% of B and remaining percentage of C. Then, the empirica formula of X is

Given atomic masses of A= 64: B=40;C = 32)

(a) AB_2C_2 (b) ABC_4 (c) A_2BC_2 (d) ABC_3

95. The rate of a reaction quadruples when temperature changes from 27°C to 57°C. Calculate the energy of activation. Given R= 8.314 K mol^{-1}, log 4 = 0.6021

(a) 3.80 kJ/mol (b) 3804 kJ/mp (c) 38.04 kJ/mol (d) 380.4 kJ/mol

96. The plot of osmotic pressure (II) vs concentration (mol L^{-1}) for a solution gives a straight line with slope 25.73 L bar-mol^{-1}. The temperature at which the osmotic pressure measurement is done is: (Use R = 0.083 bar mol^{-1} K^{-1})

(a) 25.73°C 6 (b) 12.05°C (c) 37°C (d) 310°C

97. During the preparation of Mohr's salt solution (Ferrous ammonium sulphate), which of the following acid is added to prevent hydrolysis of Fe^{2+} ion?

(a) dilute nitric acid (b) dilute sulphuric acid
(c) dilute hydrochloric acid (d) concentrated sulphuric acid

98. Mass in grams of copper deposited by passing 9.6487 A current through a voltmeter containing copper sulphate solution for 100 seconds is: (Given: Molar mass of Cu: 63 g mol^{-1}, IF= 96487 C)

(a) 31.5 g (b) 0.0315 (c) 3.15 g (d) 0.315 g

99. Identify the correct answer.

(a) Dipole moment of NF_3 is greater than that of NH_3.
(b) Three canonical forms can be drawn for CO_3^{2-} ion.
(c) Three resonance structures can be drawn for ozone.
(d) BF_3 has non-zero dipole moment.

100. Given below are two statements:

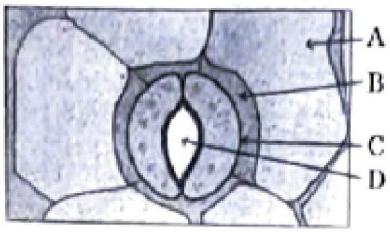
Statement 1: $\left[Co(NH_3)_6\right]^{3+}$ is a homoleptic complex whereas $\left[Co(NH_3)_4Cl_2\right]^{+}$ is a heteroleptic complex.

Statement II: Complex $\left[Co(NH_3)_6\right]^{3+}$ has only one kind of ligands but $\left[Co(NH_3)_4Cl_2\right]^{+}$ has more than one kind of ligands.

In the light of the above statements, choose the correct answer from the options given below:
(a) Statement I is true but Statement II is false. (b) Statement I is false but Statement Il is true.
(c) Both Statement I and Statement II are true. (d) Both Statement I and Statement II are false.

SECTION-A (Q.101 TO Q.135)

101. In the given figure, which component has thin outer walls and highly thickened inner walls?



(a) A (b) B (c) C (d) D

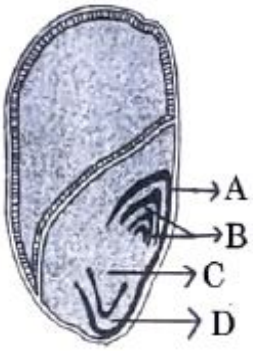
102. A transcription unit in DNA is defined primarily by the three regions in DNA and these are with respect to upstream and down stream end;
(a) Inducer, Repressor, Structural gene (b) Promotor, Structural gene, Terminator
(c) Repressor, Operator gene, Structural gene (d) Structural gene, Transposons, Operator gene

103. The equation of Verhulst-Pearl logistic growth is

$$\frac{dN}{dt} = rN\left[\frac{K-N}{K}\right]$$

From this equation, K indicates:
(a) Carrying capacity (b) Population density
(c) Intrinsic rate of natural increase (d) Biotic potential

104. Identify the part of the seed from the given figure which is destined to form root when the seed germinates.



(a) C (b) D (c) A (d) B

105. Inhibition of Succinic dehydrogenase enzyme by malonate is a classical example of
(a) Competitive inhibition (b) Enzyme activation
(c) Cofactor inhibition (d) Feedback inhibition

106. A pink flowered Snapdragon plant was crossed with a red flowered Snapdragon plant. What type of phenotype/s is/are expected in the progeny?
(a) Only pink flowered plants
(b) Red, Pink as well as white flowered plants
(c) Only red flowered plants
(d) Red flowered as well as pink flowered plants

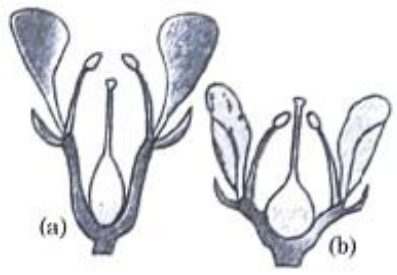
107. The type of conservation in which the threatened species are taken out from their natural habitat and placed in special setting where they can be protected and given spécial care is called;
(a) Semi-conservative method (b) Sustainable development
(c) *in-situ* conservation (d) Biodiversity conservation

108. These are regarded as major causes of biodiversity loss:
A. Over exploitation B. Co-extinction
C. Mutation D. Habitat loss and fragmentation E. Migration
Choose the correct option:
(a) A, B and E only (b) A, B and D only
(c) A, C and D only (d) A, B, C and D only

109. Which of the following are required for the dark reaction of photosynthesis?
A. Light B. Chlorophyll
C. CO_2 D. ATP E. NADPH
Choose the correct answer from the options given below:
(a) C, D and E only (b) D and E only
(c) A, B and C only (d) B, C and D only

110. Bulliform cells are responsible for
(a) Increased photosynthesis in monocots. (b) Providing large spaces for storage of sugars
(c) Inward curling of leaves in monocots (d) Protecting the plant from salt stress.

111. Identify the type of flowers based on the position of calyx, corolla and androecium with respect to the ovary from the given figures (a) and (b)



(a) (a) Perigynous; (b) Epigynous (b) (a) Perigynous; (b) Perigynous
(c) (a) Epigynous; (b) Hypogynous (d) (a) Hypogynous; (b) Epigynous

112. Which one of the following is <u>not</u> a criterion for classification of fungi?

(a) Mode of spore formation (b) Fruiting body
(c) Morphology of mycelium (d) Mode of nutrition

113. Hind II always cuts DNA molecules at a particular point called recognition sequence and it consists of:
(a) 4 bp (b) 10 bp (c) 8 bp (d) 6 bp

114. Auxin is used by gardeners to prepare weed-free lawns. But no damage is caused to grass as auxin
(a) does not affect mature monocotyledonous plants.
(b) can help in cell division in grasses, to produce growth.
(c) promotes apical dominance.
(d) promotes abscission of mature leaves only.

115. Match List I with List II

List I	**List II**
A. Two or more alternative forms of a gene	I. Back cross
B. Cross of F_1 progeny with homozygous recessive parent	II. Ploidy
C. Cross of F_1 progeny with any of the parents	III. Allele
D. Number of chromosome sets in plant	IV. Test cross

Choose the correct answer from the options given below:
(a) A-III, B-IV, C-I, D-II (b) A-IV, B-III, C-II, D-I
(c) A-I, B-II, C-III, D-IV (d) A-II, B-I, C-III, D-IV

116. Spindle fibers attach to kinetochores of chromosomes during
(a) Anaphase (b) Telophase (c) Prophase (d) Metaphase

117. Match List I with List II

List I	**List II**
A. *Clostridium butylicum*	I. Ethanol
B. *Saccharomyces cerevisiae*	II. Streptokinase
C. *Trichoderma polysporum*	III. Butyric acid
D. *Streptococcus sp.*	IV. Cyclosporin-A

Choose the correct answer from the options given below:
(a) A-III, B-I, C-IV, D-II (b) A-IV, B-I, C-III, D-II
(c) A-III, B-I, C-II, D-IV (d) A-II, B-IV, C-III, D-I

118. Which one of the following can be explained on the basis of Mendel's Law of Dominance?
A. Out of one pair of factors one is dominant and the other is recessive.
B. Alleles do not show any expression and both the characters appear as such in F_2 generation.
C. Factors occur in pairs in normal diploid plants.
D. The discrete unit controlling a particular character is called factor.
E. The expression of only one of the parental characters is found in a monohybrid cross.
Choose the correct answer from the options given below:
(a) B, C and D only (b) A, B, C, D and E
(c) A, B and C only (d) A, C, D and E only

119. What is the fate of a piece of DNA carrying only gene of interest which is transferred into an alien organism?
A. The piece of DNA would be able to multiply itself independently in the progeny cells of the organism.
B. It may get integrated into the genome of the recipient.
C. It may multiply and be inherited along with the host DNA.

D. The alien piece of DNA is not an integral part of chromosome.

E. It shows ability to replicate.

Choose the correct answer from the options given below:

(a) B and C only (b) A and E only (c) A and B only (d) D and E only

120. How many molecules of ATP and NADPH are required for every molecule of CO_2 fixed in the Calvin cycle?

(a) 3 molecules of ATP and 3 molecules of NADPH

(b) 3 molecules of ATP and 2 molecules of NADPH

(c) 2 molecules of ATP and 3 molecules of NADPH

(d) 2 molecules of ATP and 2 molecules of NADPH

121. In a plant, black seed color (BB/Bb) is dominant over white seed color (bb). In order to find out the genotype of the black seed plant, with which of the following genotype will you cross it?

(a) Bb (b) BB/Bb (c) BB (d) bb

122. Lecithin, a small molecular weight organic compound found in living tissues, is an example of

(a) Glycerides (b) Carbohydrates (c) Amino acids (d) Phospholipids

123. Match List I with List II

List I	**List II**
A. *Rhizopus*	I. Mushroom
B. *Ustilago*	II. Smut fungus
C. *Puccinia*	III. Bread mould
D. *Agaricus*	IV. Rust fungus

Choose the correct answer from the options given below:

(a) A-III, B-II, C-I, D-IV (b) A-IV, B-III, C-II, D-I

(c) A-III, B-II, C-IV, D-I (d) A-I, B-III, C-II, D-IV

124. Tropical regions show greatest level of species richness because

A. Tropical latitudes have remained relatively undisturbed for millions of years, hence more time was available for species diversification.

B. Tropical environments are more seasonal.

C. More solar energy is available in tropics.

D. Constant environments promote niche specialization.

E. Tropical environments are constant and predictable.

Choose the correct answer from the options given below:

(a) A, B and E only (b) A, B and D only (c) A, C, D and E only (d) A and B only

125. Match List I with List II

List I	**List II**
A. Nucleolus	I. Site of formation of glycolipid
B. Centriole	II. Organization like the cartwheel
C. Leucoplasts	III. Site for active ribosomal RNA synthesis
D. Golgi apparatus	IV. For storing nutrients

Choose the correct answer from the options given below:

(a) A-III, B-IV, C-II, D-I (b) A-I, B-II, C-III, D-IV

(c) A-III, B-II, C-IV, D-I (d) A-II, B-III, C-I, D-IV

126. The lactose present in the growth medium of bacteria is transported to the cell by the action of

(a) Permease (b) Polymerase (c) Beta-galactosidase (d) Acetylase

127. Given below are two statements:

Statement I: Chromosomes become gradually visible under light microscope during leptotene stage.

Statement II: The begining of diplotene stage is recognized by dissolution of synaptonemal complex.

In the light of the above statements, choose the correct answer from the options given below:

(a) Statement I is true but Statement II is false (b) Statement I is false but Statement II is true

(c) Both Statement I and Statement II are true (d) Both Statement I and Statement II are false

128. Formation of interfascicular cambium from fully developed parenchyma cells is an example for

(a) Dedifferentiation (b) Maturation (c) Differentiation (d) Redifferentiation

129. Given below are two statements:

Statement I: Parenchyma is living but collenchyma is dead tissue.

Statement II: Gymnosperms lack xylem vessels but presence of xylem vessels is the characteristic of angiosperms.

In the light of the above statements, choose the correct answer from the options given below:

(a) Statement I is true but Statement II is false (b) Statement I is false but Statement II is true

(c) Both Statement I and Statement II are true (d) Both Statement I and Statement II are false

130. Identify the set of correct statements:

A. The flowers of *Vallisneria* are colourful and produce nectar.

B. The flowers of waterlily are not pollinated by water.

C. In most of Water-pollinated species, the pollen grains are protected from wetting.

D. Pollen grains of some hydrophytes are long and ribbon like.

E. In some hydrophytes, the pollen grains are carried passively inside water.

Choose the correct answer from the options given below:

(a) A, C, D and E only (b) B, C, D and E only

(c) C, D and E only (d) A, B, C and D only

131. Which of the following is an example of actinomorphic flower?

(a) Pisum (b) Sesbania (c) Datura (d) Cassia

132. The capacity to generate a whole plant from any cell of the plant is called:

(a) Differentiation (b) Somatic hybridization

(c) Totipotency (d) Micropropagation

133. Given below are two statements:

Statement I: Bt toxins are insect group specific and coded by a gene *cry* IAc.

Statement II: Bt toxin exists as inactive protoxin in *B. thuringiensis*. However, after ingestion by the insect the inactive protoxin gets converted into active form due to acidic pH of the insect gut.

In the light of the above statements, choose the correct answer from the options given below:

(a) Statement I is true but Statement II is false

(b) Statement I is false but Statement II is true

(c) Both Statement I and Statement II are true

(d) Both Statement I and Statement II are false

134. List of endangered species was released by

(a) FOAM (b) IUCN (c) GEAC (d) WWF

135. The cofactor of the enzyme carboxypeptidase is:

(a) Flavin (b) Haem (c) Zinc (d) Niacin

SECTION-B (Q.136 TO Q.150)

136. Match List I with List II

List I	List II
A. Citric acid cycle	I. Cytoplasm
B. Glycolysis	II. Mitochondrial matrix
C. Electron transport system	III. Intermembrane space of mitochondria
D. Proton gradient	IV. Inner mitochondrial membrane

Choose the correct answer from the options given below:

(a) A-III, B-IV, C-I, D-II (b) A-IV, B-III, C-II, D-I

(c) A-I, B-II, C-III, D-IV (d) A-II, B-I, C-IV, D-III

137. Which of the following statement is correct regarding the process of replication in E.coli?

(a) The DNA dependent DNA polymerase catalyses polymerization in 5'$\rightarrow$3' as well as 3'$\rightarrow$5' direction.

(b) The DNA dependent DNA polymerase catalyses polymerization in 5'$\rightarrow$3' direction.

(c) The DNA dependent DNA polymerase catalyses polymerization in one direction that is 3'$\rightarrow$5'.

(d) The DNA dependent RNA polymerase catalyses polymerization in one direction, that is 5'$\rightarrow$3'.

138. Match List I with List II

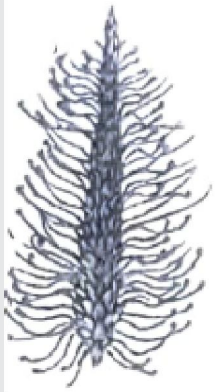
List I	List II
A. Robert May	I. Species-Area relationship
B. Alexander von Humboldt	II. Long term ecosystem experiment using out door plots
C. Paul Ehrlich	III. Global species diversity at about 7 million
D. David Tilman	IV. Rivet popper hypothesis

Choose the correct answer from the options given below:

(a) A-I, B-III, C-II, D-IV (b) A-III, B-IV, C-II, D-I

(c) A-II, B-III, C-I, D-IV (d) A-III, B-I, C-IV, D-II

139. Identify the correct description about the give figure:



(a) Cleistogamous flowers showing autogamy.

(b) Compact inflorescence showing complete autogamy.

(c) Wind pollinated plant inflorescence showing flowers with well exposed stamens.

(d) Water pollinated flowers showing stamens with mucilaginous covering.

140. Identify the step in tricarboxylic acid cycle, which does not involve oxidation of substrate.

(a) Succinyl-CoA $\rightarrow$ Succinic acid (b) Isocitrate $\rightarrow$ α -ketoglutaric acid

(c) Malic acid $\rightarrow$ Oxaloacetic acid (d) Succinic acid $\rightarrow$ Malic acid

141. Given below are two statements:

Statement I: In C_3 plants, some O_2 binds to RuBisCO, hence CO_2 fixation is decreased.

Statement II: In C_4 plants, mesophyll cells show very little photorespiration while bundle sheath cells do not show photorespiration.

In the light of the above statements, choose the correct answer from the options given below:
(a) Statement I is true but Statement II is false
(b) Statement I is false but Statement II is true
(c) Both Statement I and Statement II are true
(d) Both Statement I and Statement II are false

142. In an ecosystem if the Net Primary Productivity (NPP) of first trophic level is $100x\left(kcal\ m^{-2}\right)yr^{-1}$, what would be the GPP (Gross Primary Productivity) of the third trophic level of the same ecosystem?

(a) $10x\left(kcal\ m^{-2}\right)yr^{-1}$ (b) $\frac{100x}{3x}\left(kcal\ m^{-2}\right)yr^{-1}$

(c) $\frac{x}{10}\left(kcal\ m^{-2}\right)yr^{-1}$ (d) $x\left(kcal\ m^{-2}\right)yr^{-1}$

143. Match List I with List II

List I	**List II**
A. GLUT-4	1. Hormone
B. Insulin	II. Enzyme
C. Trypsin	III. Intercellular ground substance
D. Collagen	IV. Enables glucose transport into cells

Choose the correct answer from the options given below:
(a) A-II, B-III, C-IV, D-I (b) A-III, B-IV, C-I, D-II
(c) A-IV, B-I, C-II, D-III (d) A-I, B-II, C-III, D-IV

144. Match List I with List II

List I	**List II**
A. Frederick Griffith	I. Genetic code
B. Francois Jacob & Jacque Monod	II. Semi-conservative mode of DNA replication
C.Har Gobind Khorana	III. Transformation
D. Meselson & Stahl	IV. Lac operon

Choose the correct answer from the options given below:
(a) A-II, B-III, C-IV, D-I (b) A-IV, B-I, C-II, D-III
(c) A-III, B-II, C-I, D-IV (d) A-III, B-IV, C-I, D-II

145. The DNA present in chloroplast is:
(a) Linear, single stranded (b) Circular, single stranded
(c) Linear, double stranded (d) Circular, double stranded

146. Spraying sugarcane crop with which of the following plant growth regulators, increases the length of stem, thus, increasing the yield?
(a) Cytokinin (b) Abscisic acid (c) Auxin (d) Gibberellin

147. Read the following statements and choose the set of correct statements:
In the members of Phaeophyceae,
A. Asexual reproduction occurs usually by biflagellate zoospores.
B. Sexual reproduction is by oogamous method only.
C. Stored food is in the form of carbohydrates which is either mannitol or laminarin.
D. The major pigments found are chlorophyll a, c and carotenoids and xanthophyll.

E. Vegetative cells have a cellulosic wall, usually covered on the outside by gelatinous coating of algin.

Choose the correct answer from the options given below:

(a) A, C, D and E only (b) A, B, C and E only

(c) A, B, C and D only (d) B, C, D and E only

148. Which of the following are fused in somatic hybridization involving two varieties of plants?

(a) Protoplasts (b) Pollens (c) Callus (d) Somatic embryos

149. Match List I with List II

List I	**List II**
A. Rose	I. Twisted aestivation
B. Pea	II. Perigynous flower
C. Cotton	III. Drupe
D. Mango	IV. Marginal placentation

Choose the correct answer from the options given below:

(a) A-IV, B-III, C-II, D-I (b) A-II, B-III, C-IV, D-I

(c) A-II, B-IV, C-I, D-III (d) A-I, B-II, C-III, D-IV

150. Match List I with List II

List I (Types of Stamens)	**List II (Example)**
A. Monoadelphous	I. Citrus
B. Diadelphous	II. Pea
C. Polyadelphous	III. Lily
D. Epiphyllous	IV. China-rose

Choose the correct answer from the options given below: ?

(a) A-I, B-II, C-IV, D-III (b) A-III, B-I, C-IV, D-II

(c) A-IV, B-II, C-I, D-III (d) A-IV, B-I, C-II, D-III